

Material that prevents radioactive emission from passing through it	Number of radioactive emissions/counts per second - also known as "rate of decay"	Release of alpha particle from unstable nucleus (a 2 <sup>+</sup> helium ion or a helium nucleus)	Number of protons in a nucleus
<b>Absorber</b>	<b>Activity</b>	<b>Alpha emission</b>	<b>Atomic number</b>
Natural radiation from everyday surroundings	The way charged particles are forced to move in a magnetic field	Release of beta particle from unstable nucleus (a single electron)	The number of counts on a Geiger counter (each count is caused by radiation causing ionisation)
<b>Background radiation</b>	<b>Behaviour in a magnetic field</b>	<b>Beta emission</b>	<b>Count rate</b>
Atoms are the smallest building blocks of matter and are indivisible (can't be split up)	Particles close together	Negatively charged subatomic particle	Release of gamma wave from unstable nucleus (Electromagnetic spectrum)
<b>Dalton's model of the atom</b>	<b>Dense</b>	<b>Electron</b>	<b>Gamma emission</b>
Device to measure radioactivity	Rutherford experiment where he bombarded a thin gold foil with alpha particles in a vacuum	Time taken for half the radioactive nuclei in a sample to decay	Remove electrons from orbit by collision
<b>Geiger counter</b>	<b>Gold foil experiment</b>	<b>Half life</b>	<b>Ionisation</b>

A measure of the ability to remove electrons	Same number of protons, different number of neutrons or same element but different number of neutrons	A sheet of lead used to absorb alpha and beta and limit gamma	Number of protons and neutrons
<b>Ionising ability</b>	<b>Isotope</b>	<b>Lead screen</b>	<b>Mass number</b>
Charge that attracts positive charges and repels negative charges	Uncharged subatomic particle	An equation to show a radioactive change with balanced atomic and mass numbers	Small, dense centre of atom containing neutrons and protons
<b>Negative</b>	<b>Neutron</b>	<b>Nuclear equation</b>	<b>Nucleus</b>
Ability of radioactive emissions to pass through materials	A neutral atom of positive charge with negative charges embedded into it	Charge that repels positive charges and attracts negative charges	Positively charged subatomic particle
<b>Penetration ability</b>	<b>Plum pudding model</b>	<b>Positive</b>	<b>Proton</b>
Emission of $\alpha$ , $\beta$ or $\gamma$ in order to transmute (change contents of nucleus) or to lose excess energy	Source (point of origin) of radioactive emissions	Number of radioactive emissions/counts per second - also known as "activity"	Neutrons and protons in a small dense nucleus with orbiting electrons and large gaps within the atom
<b>Radioactive decay</b>	<b>Radioactive source</b>	<b>Rate of decay</b>	<b>Rutherford's model of the atom</b>

<p>The way the model suggests the atom is arranged (Dalton, Thomson, Rutherford)</p>	<p>The plum pudding model</p> 	<p>Alpha, beta or gamma</p>	<p>Nucleus that will emit radiation</p>
<p><b>Structure of the atom</b></p>	<p><b>Thomson's model of the atom</b></p>	<p><b>Type of radiation</b></p>	<p><b>Unstable nuclei</b></p>
<p>Chamber with no particles inside it</p>			
<p><b>Vacuum chamber</b></p>			